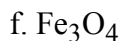
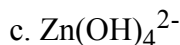
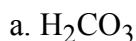


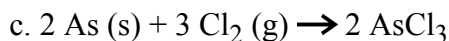
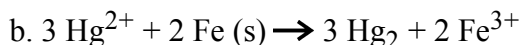
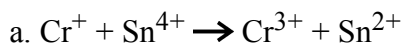
# Practice Problems: Redox Reactions

1. Determine the oxidation number of the elements in each of the following compounds:



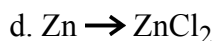
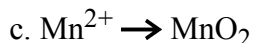
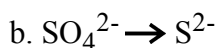
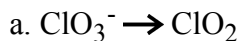
[Hint](#)

2. Identify the species being oxidized and reduced in each of the following reactions:



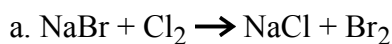
[Hint](#)

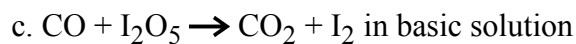
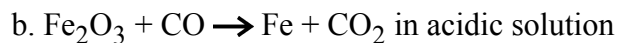
3. Would you use an oxidizing agent or reducing agent in order for the following reactions to occur?



[Hint](#)

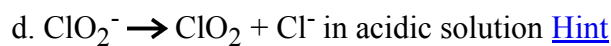
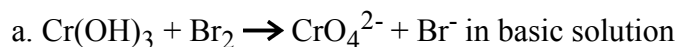
4. Write balance equations for the following redox reactions:



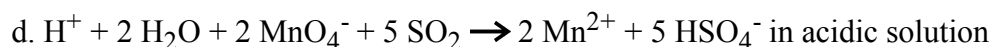
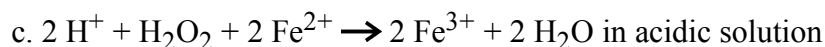
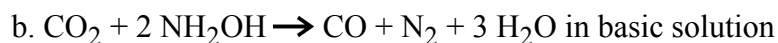


[Hint](#)

5. Write balanced equations for the following reactions: [Hint](#)



6. Write the balanced half reactions of the following reactions:



[Hint](#)

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